

Progress Report on Collaborative Research on the Development of Chelators for a Surface Decontamination Foam

M. Sutton

July 15, 2009

Disclaimer

This document was prepared as an account of work sponsored by an agency of the United States government. Neither the United States government nor Lawrence Livermore National Security, LLC, nor any of their employees makes any warranty, expressed or implied, or assumes any legal liability or responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by trade name, trademark, manufacturer, or otherwise does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States government or Lawrence Livermore National Security, LLC. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States government or Lawrence Livermore National Security, LLC, and shall not be used for advertising or product endorsement purposes.

This work performed under the auspices of the U.S. Department of Energy by Lawrence Livermore National Laboratory under Contract DE-AC52-07NA27344.

Lawrence Livermore National Laboratory

Progress Report on Collaborative Research on the Development of Chelators for a Surface Decontamination Foam - June 30th, 2009 -



Mark Sutton, PhD.

Chemical Science Division

Cs, Sr & Co – Chelate Models

Running the models:

Chemical thermodynamic modeling using data from the NIST Critical Stability Constant Database v46.

No data available for Cs DTPA and Cs NTA.

Not kinetic – assumes complete reaction.

1 micromolar metal solutions.

1 micromolar to 1 molar chelator solution.

Titrate pH.

Some models did not converge at high chelate concentration.

Cs, Sr & Co – Chelate Models

Data Caveats:

Data show optimum pH and chelate concentration for maximum chelator efficiency.

Models do not show speciation....

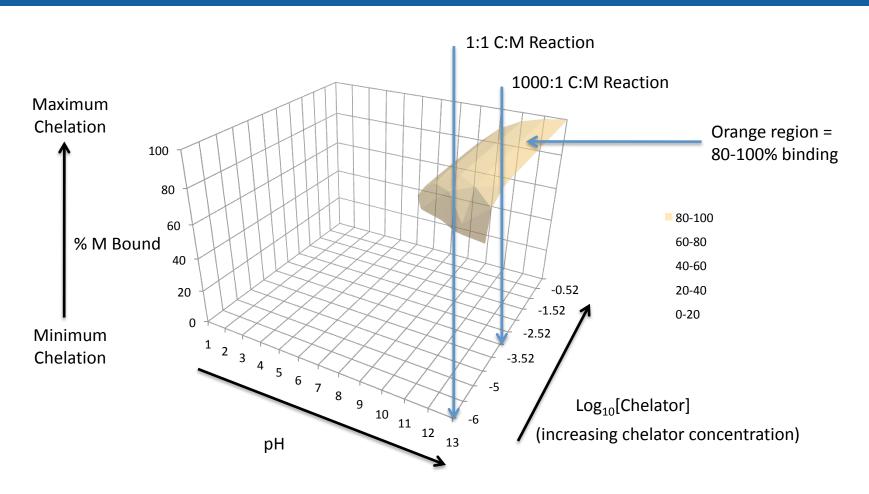
EDTA binds lots of metals effectively (i.e. non-specific).

EDTA shows >60% efficiency for Cs at high pH, but in the real world (with natural background levels of other metals), EDTA will never bind Cs efficiently.

However, it is possible to vary pH to alter EDTA specificity for Co in a mixture of Sr and Cs....

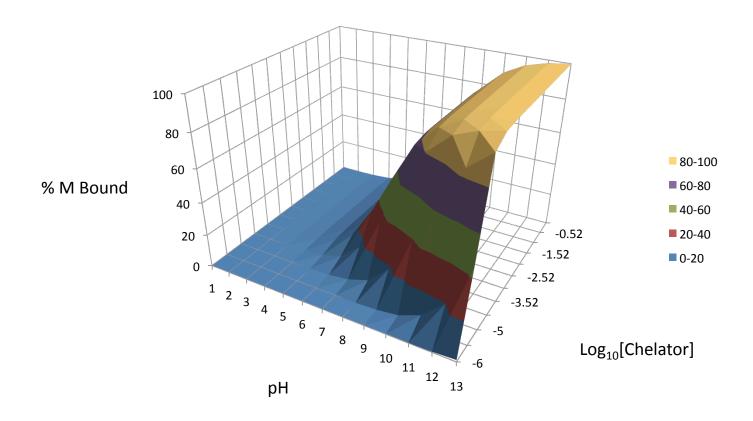
EDTA binds to Co at pH 2-12, Sr at pH 4-12 and poorly to Cs. Therefore, to selectively bind Co, maintain pH <4.

M – Chelate Contours

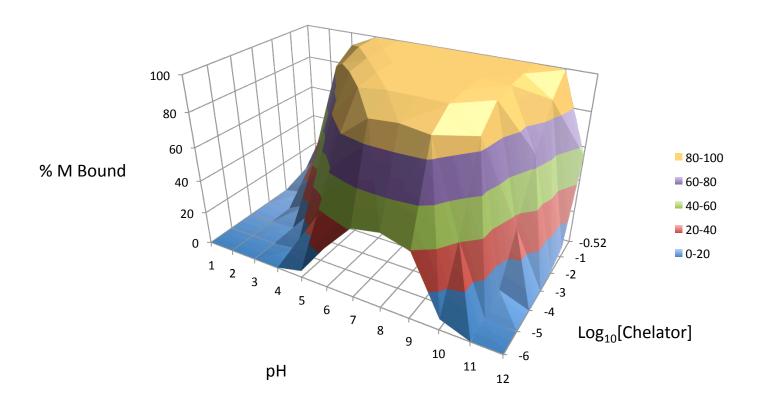




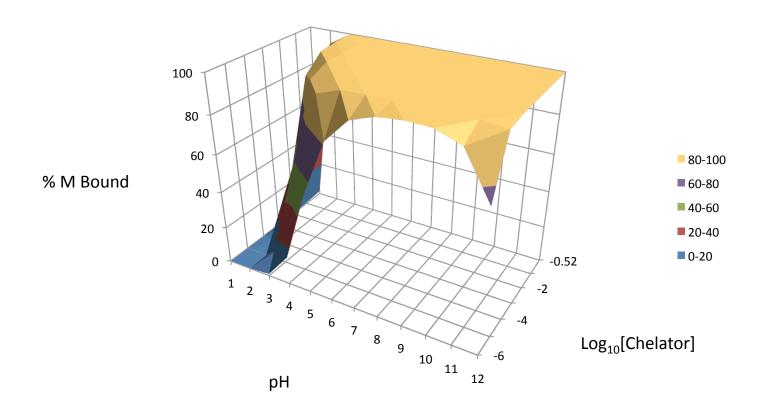
Cs – Prussian Blue



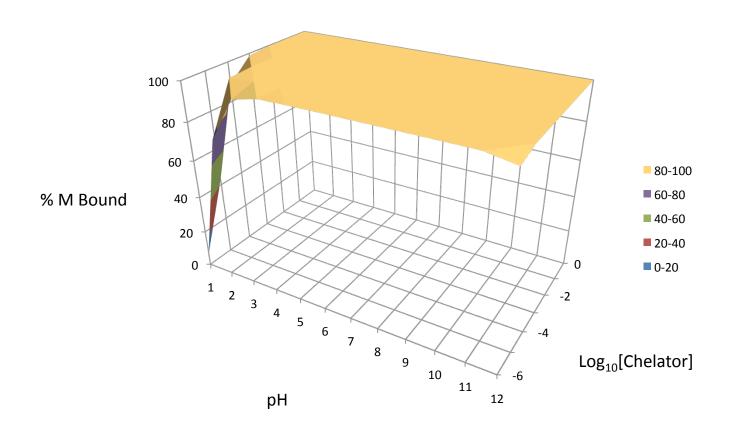
Co – Citrate



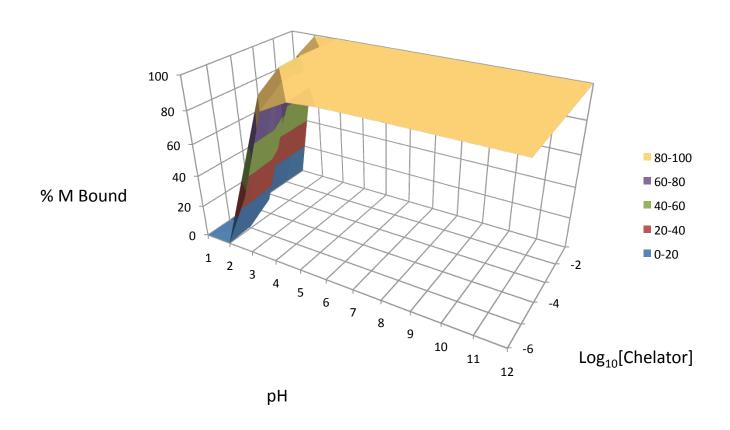
Co - NTA



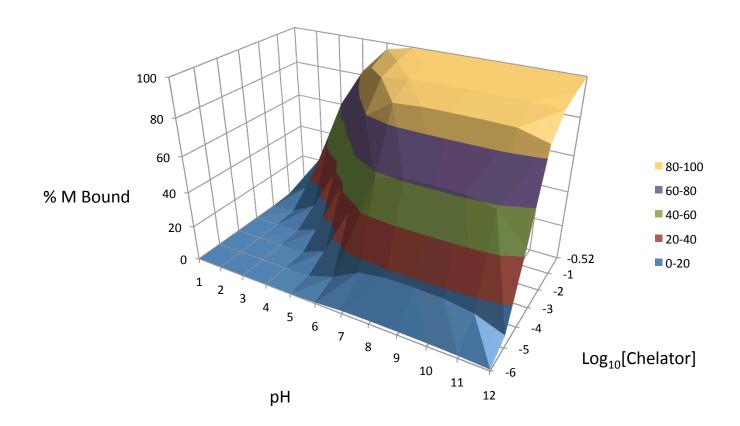
Co - EDTA



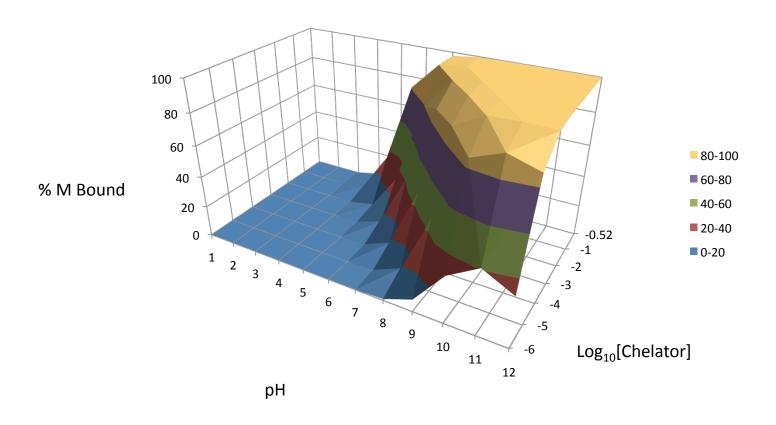
Co - DTPA



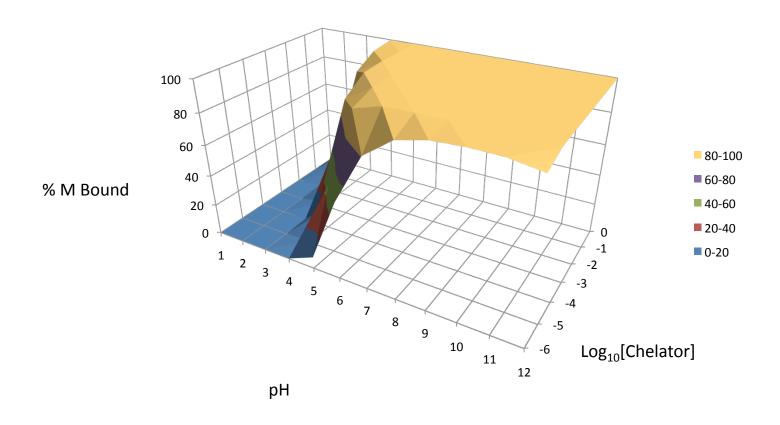
Sr – Citrate



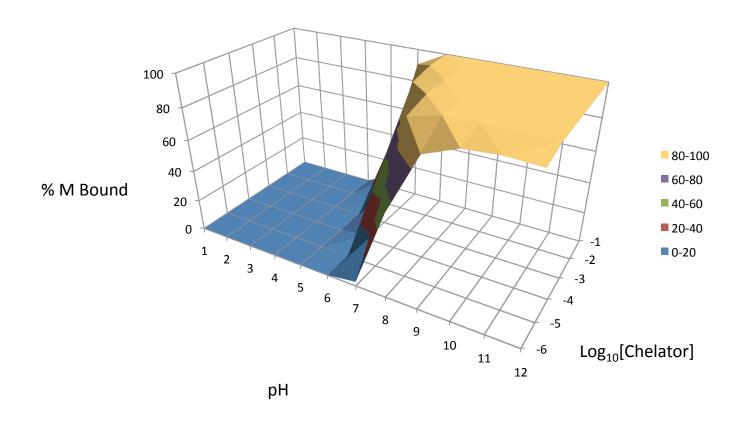
Sr - NTA



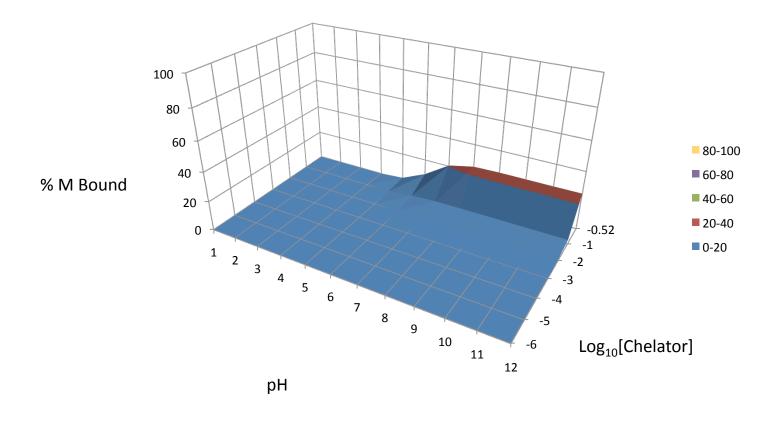
Sr – EDTA



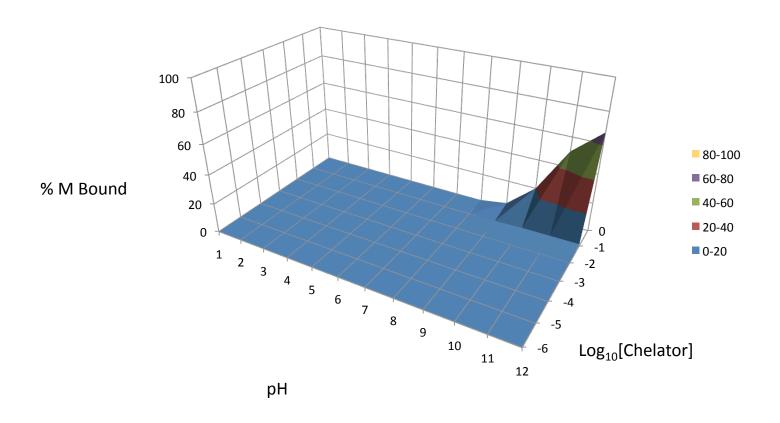
Sr - DTPA



Cs - Citrate



Cs – EDTA



Disclaimer

This document was prepared as an account of work sponsored by an agency of the United States government. Neither the United States government nor Lawrence Livermore National Security, LLC, nor any of their employees makes any warranty, expressed or implied, or assumes any legal liability or responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by trade name, trademark, manufacturer, or otherwise does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States government or Lawrence Livermore National Security, LLC. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States government or Lawrence Livermore National Security, LLC, and shall not be used for advertising or product endorsement purposes.